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BALANCING EQUATIONS

W 301 Balancing Equations Worksheet1) $\text{H}_3\text{PO}_4 + \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$ 2) $\text{K} + \text{B}_2\text{O}_3 \rightarrow \text{K}_2\text{O} + \text{B}$ 3) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ 4) $\text{Na} + \text{NaNO}_3 \rightarrow \text{Na}_2\text{O} + \text{N}_2$ 5) $\text{C} + \text{S}_8 \rightarrow \text{CS}_2$ 6) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$ 7) $\text{N}_2 + \text{O}_2 \rightarrow \text{N}_2\text{O}_5$ 8) $\text{H}_3\text{PO}_4 + \text{Mg}(\text{OH})_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$ 9) $\text{NaOH} + \text{H}_2\text{CO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$ 10) $\text{KOH} + \text{HBr} \rightarrow \text{KBr} + \text{H}_2\text{O}$ 11) $\text{Na} + \text{O}_2 \rightarrow \text{Na}_2\text{O}$ 12) $\text{Al}(\text{OH})_3 + \text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + \text{H}_2\text{O}$ 13) $\text{Al} + \text{S}_8 \rightarrow \text{Al}_2\text{S}_3$ 14) $\text{Cs} + \text{N}_2 \rightarrow \text{Cs}_3\text{N}$ 15) $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$ 16) $\text{Rb} + \text{RbNO}_3 \rightarrow \text{Rb}_2\text{O} + \text{N}_2$ 17) $\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2$...

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BALANCING EQUATIONS - CLASS PRACTICE

Directions: You MUST show your work for each problem. List metals, non-metals, then H and O. Balancing is easier when using this order. Check your answers for accuracy. Once you think you have the hang of it you will then pick up your homework to turn in tomorrow. 1. $\text{NaOH} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{NaNO}_3 + \text{Pb}(\text{OH})_2$ 2. $\text{CaSO}_4 + \text{LiF} \rightarrow \text{CaF}_2 + \text{Li}_2\text{SO}_4$ 3. $\text{AlF}_3 + \text{Na}_2\text{CO}_3 \rightarrow \text{NaF} + \text{Al}_2(\text{CO}_3)_3$ 4. $\text{Ag} + \text{CuSO}_4 \rightarrow \text{Ag}_2\text{SO}_4 + \text{Cu}$ 5. $\text{C} + \dots$

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Balancing Chemical Equations Here is an example equation $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ It is an unbalanced equation sometimes also called a skeleton equation This means that there are UNEQUAL numbers at least one atom on each side of the arrow In the example equation there are two atoms of hydrogen on each side BUT there are two atoms of oxygen on the left side and only one on the right side Remember this A balanced equati...

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